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In water or a water solution, the solution is acidic if the hydrogen ion concentration is greater than the hydroxide (OH-) ion concentration is neutral when the concentration, and the solution is desic if the hydroxide ion concentration is greater than the hydroxide (OH-) ion concentration is greater than the hydroxide (OH-) ion concentration is greater than the hydroxide (OH-) ion concentration is greater than the hydroxide ion concentration is greater than properties of basic solutions are due to the high concentration of ...

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Consider an acetic acid solution at equilibrium. (1) H2O(I) + CH3CO2H(aq) H3O+(aq) + CH3CO2-(aq) There are three quantities commonly measured and/or known for a weak acid solution: the value of the equilibrium constant from reference sources, the solution pH from experimental measurement, and.

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If this acid is 2.0% dissociated at equilibrium, 0.10 x 0.020 = 0.0020 moles of acid per liter of solution are ionized. Thus acids may be further classified as strong bases and weak bases. For example, the -COOH group is referred to as a carboxylic acid. The concentration of the undissociated acid, HX ...

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Bases are substances that produce hydroxide ions (OH-). Soluble bases are called alkalis and the solutions. Acids taste sour, while bases taste bitter. Common laboratory acids include hydrochloric acid (HCI), sulfuric acid (H2SO4) and nitric acid (HNO3).

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